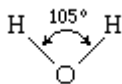
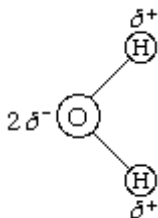


The water molecule is composed of two hydrogen atoms, each linked by a single chemical bond to an oxygen atom.

The water molecule is not linear but bent in a special way. The two hydrogen atoms are bound to the oxygen atom at an angle of  $104.5^\circ$ .



Because an oxygen atom has a greater electronegativity than a hydrogen atom, the O–H bonds in the water molecule are polar, with the oxygen bearing a partial negative charge and the hydrogens having a partial positive charge.



Hydrogen atoms in water molecules are attracted to regions of high electron density and can form weak linkages, called hydrogen bonds, with those regions. This means that the hydrogen atoms in one water molecule are attracted to the nonbonding electron pairs of the oxygen atom on an adjacent water molecule. The structure of liquid water is believed to consist of aggregates of water molecules that form and reform continually.